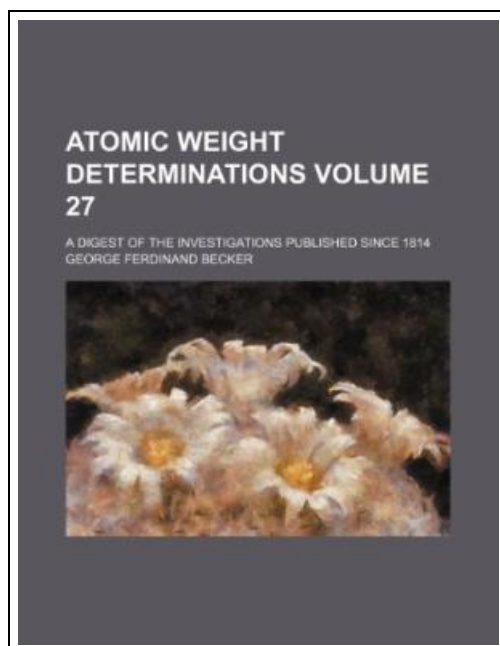


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1883 Excerpt: . c.) L. J. Thenard: F. H. Wollaston: 55.2 (O = 16); 345 (O = 100). Thenard determined the composition of the oxide at 22.5 O and 77.5 Fe, whence Wollaston calculates the value. (Phil. Trans., 104, 1814, 21.) J. J. Berzelius: 54.27 (O = 16); 339.213 (O = 100). Determined by repeated experiments on the oxidation of iron, such as is used for piano wire, with nitric acid. The carbon was determined and allowed for. Berzelius in his Lehrbuch shows that the error in this determination lay in the unsuspected presence of soluble silica and on reanalysis he found enough of it to correct the number when taken into account. (Poggend. Ann., 8, 1826, 185.) G. Magnus: 54.25 (O = 16); 339.06 (O = 100). Magnus experiments were made by reducing ferric oxide in a current of hydrogen at about the temperature of boiling mercury. He regarded them simply as confirmatory of Berzelius number. Poggend. Ann., 3, 1825, 84.) F. Stromeyer: 55.6 (O = 16). Determined by reducing ferric oxide at a red heat by hydrogen. The oxide is reduced only with great difiiculty at a lower temperature. The mean of the experiments gave the oxygen contents at 30.15 per cent., whence I have calculated the value. (Poggend. Ann., 6, 1826, 475.) H. Capitaine: 51.36 (O = 16); 321 (O = 100). Determined by the peroxidation of galvanically reduced iron and by measuring the hydrogen evolved on the solution of iron in sulphuric acid. (Annul....



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