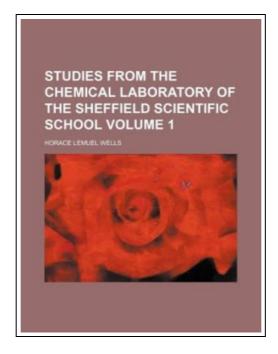
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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*\*This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1901 Excerpt: .chromium alums. With fa and fa, equal to zero, the solubilities of the vanadium alums are obtained, and by eliminating the second term as well, those of the aluminium salts. We thus have a general solubility equation, all the terms of which can be referred to two variables; the one of these, 6, applying to the one class of elements in the compounds, the other, k, to the second class. I believe that aside from the instance of addition properties, which the solubilities certainly are not, this is the first case in which a mathematical relation between the corresponding properties of a class of compounds has been found capable of expression. Below I give the values of 6 and p for each of the alkali metals, and the solubilities as calculated by the general formula. With the latter, the column D contains the difference between the calculated and observed solubilities, and D the variation from the values serving as a basis for the derivation of the constants. - p Cs. 2 6 5 10.945 Eb. 27 11 19 12.254 Tl. 57 56 4 20.5406 NH. 73 45 5 38.943 In the following the solubilities are expressed in hundredtha of a gram-molecule: In conclusion I may state that while the data now at hand are too few for accurate analysis, such determinations as I have made at temperatures other than 25 indicate that a general solubility formula for all temperatures will not be difficult of derivation. I...



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