



A Course in Qualitative Chemical Analysis

By Charles Baskerville

Theclassics.us, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1916 edition. Excerpt: . PART I THE METALS Reactions of the Metals of Group I The metals, silver, mercury (-ous), and lead, comprising this group, are distinguished from all others by the insolubility of their chlorides in water and in dilute acids. With the exception of the nitrates and acetates, which are colorless, nearly all the salts of the metals of this group are insoluble in water. Silver 1. Hydrochloric acid or a soluble chloride, when added to solutions of silver salts, gives a white, curdy precipitate of silver chloride (AgCl) which darkens on exposure to light. The precipitate is insoluble in water, the solubility being approximately 1 part in 700,000 parts of water; it is insoluble in dilute acids and in dilute aqua regia, but is somewhat soluble in concentrated acids. Ammonium hydroxide readily dissolves it, with the formation of silver ammonia chloride [Ag DEGREESH3 DEGREESC]:...



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