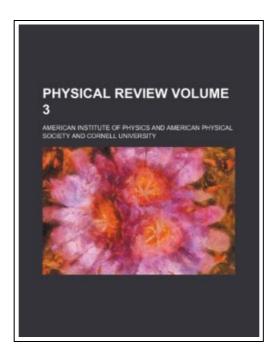
# Physical Review Volume 3



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## Reviews

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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*. This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1896 Excerpt: .has pointed out that these potential differences are functions of the metal forming the electrode and of the anion. This can hardly be accounted for on the Ostwald-Nernst hypothesis. If the potential difference between Hg and KC1 or KBr solutions are due to the amount of mercury as ion which has gone into solution, we must say that the amount varies as we change from KC1 to KBr, or, in other words, that the negative ion has an effect. This is quite apart from the difficulty of accounting for the sign of the potential difference. I do not see that the relative solubilities of mercurous chloride and bromide can be used to help out matters, because we do not have a saturated solution at all, and the difference in the electromotive forces is more likely to be connected with the difference of solubility as cause than as effect. There are no experimental data, so far as I know, on potential differences at the contact surface of reversible electrodes except some measurements by Neumann,3 and these do not establish the point they were intended to prove owing to an unfortunate choice of solutions. He measured the potential difference between thallium and solutions of thallium salts. Most of the salts were salts of organic acids, and Ostwald4 had already found that when the negative ion was an organic radical its nature was immaterial To settle this question one should take negative ions which show marked differences with non-reversible...



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