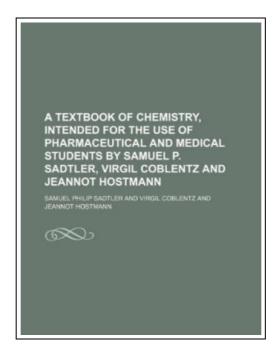
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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand ******. This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1918 Excerpt: .in water it decomposes it with some violence. The salts of strontium show great similarities to those of barium. STRONTIUM AND THE HALOGENS Strontium Chloride, SrCl2.6H20.--This salt is prepared by saturating hydrochloric acid with the native carbonate and digesting the solution for some time with excess of carbonate, by which metals like aluminum, iron, and manganese are removed. The hot concentrated solution deposits, on cooling, acicular crystals of the hexagonal system. These crystals have a sharp, bitterish taste, and deliquesce on exposure to the air. They are soluble in 1.8 parts of water at ordinary temperatures, and in 0.9 part of boiling water. The ready solubility of this salt in alcohol offers a ready means for its separation from barium chloride. On the application of heat the salt loses its water of crystallization at 100, and at 8730 it melts. The anhydrous salt is soluble in absolute alcohol. Strontium Bromide, SrBr2.6H20. Strontii Bromidum, U.S.P.--To hydrobromic acid, diluted with an equal volume of water, a slight excess over an equimolecular quantity of strontium carbonate is added in portions until the effervescence has subsided. The mixture is then digested for a few hours, filtered, washed, and the nitrate concentrated for crystallization. SrCO, + 2HBr = SrBr, + H.O + CO Strontium bromide may also be prepared by boiling a concentrated solution of a molecular quantity of ammonium bromide with an equimolecular quantity of strontium hydroxide until ammonia gas ceases to be given off. After filtration and concentration the...

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